

Chem 1- Moles Whiteboard Problems

- Find the empirical formula for a compound which contains 32.8% chromium and 67.2% chlorine. CrCl_3
- A compound is 24.7% Calcium, 1.2% Hydrogen, 14.8% Carbon, and 59.3% Oxygen. Write the empirical formula. $\text{CaH}_2\text{C}_2\text{O}_6 \rightarrow \text{Ca}(\text{HCO}_3)_2$
- A compound is 21.20% Nitrogen, 6.06% Hydrogen, 24.30% Sulfur, and 48.45% Oxygen. Write the empirical formula. $\text{N}_2\text{H}_8\text{SO}_4$
- A compound is 44.82% Potassium, 18.39% Sulfur, and 36.79% Oxygen. Write the empirical formula. ~~K_2SO_3~~ K_2SO_4
- A compound is 52.0% Zinc, 9.6% Carbon, and 38.4% Oxygen. Calculate the empirical formula of the compound. ZnCO_3
- A compound contains 21.6% Na, 33.3% Cl, and 45.1% O. Write the empirical formula. NaClO_3
- Find the percent oxygen in NaBrO_3 . 31.8
- Find the percent iron in Fe_2O_3 . 70.70
- Find the percent nitrogen in N_2O_4 . 30.470
- Find the percent silver in AgNO_3 . 63.570
- Find the percent hydrogen in $\text{C}_6\text{H}_{12}\text{O}_6$. 6.6770
- Find the molar mass of $(\text{NH}_4)_2(\text{CO}_3)$ 96
- How many molecules are in 0.65 moles of KBr ? 3.91×10^{23}
- How many moles are in 1.65×10^{23} molecules of Al_2O_3 ? 0.274
- How many grams are in 4.5 moles of ZnO ? 366.3
- How many grams are in 0.23 moles of HBr ? 18.6
- How many moles are in 58 grams of K_2CrO_4 ? 0.299
- How many moles are in 124 grams of CaCO_3 ? 1.24
- How many molecules are in 53 grams of HNO_3 ? 5.06×10^{23}
- How many grams are in 7.58×10^{23} molecules of N_2 ? 35.3
- How many grams are in 6.3×10^{22} molecules of H_2O ? 1.88